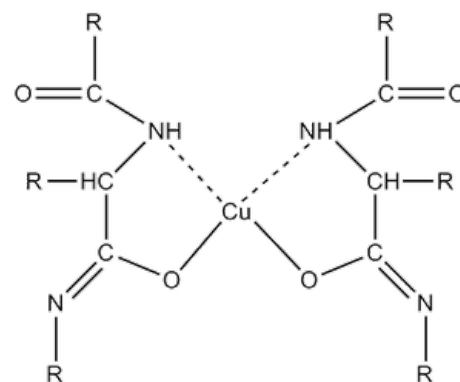


# Biuret reaction

In an **alkaline environment in the presence of copper salts** give the proteins a purple colour, suitable for photometric determination<sup>[1] [2]</sup>. In simple terms, complex compounds of  $\text{Cu}^{2+}$  with peptide bond ions are formed<sup>[3]</sup>. The formed complex strongly absorbs light in the 540-560 nm region. The intensity of the colour of the complex is measured photometrically and the absorbation is directly proportional to the concentration of protein. The biuret reaction (so-called biuret test) is generally provided by substances containing at least two peptide bonds (-CO-NH-) or two -CO-NH<sub>2</sub> groups in molecule. Thus, the reaction is not specific to proteins only. The simplest compound reacting with copper salts in an alkaline environment is **biuret** containing two peptide bonds. Amino acids and dipeptides do not react.

The biuret reagent includes **copper sulfate**, which provides  $\text{Cu}^{2+}$  for the formation of complexes with peptide bonds, and an alkalizing component (hydroxide), which converts the peptide bond to an enol form, allowing oxygen atoms to participate in the complex. Other components of the reagent are **potassium sodium tartrate**, whichas a complexing agent prevents precipitation  $\text{Cu}^{2+}$  to  $\text{Cu}(\text{OH})_2$ , and **potassium iodide**, which protects  $\text{Cu}^{2+}$  from autoreduction. The sensitivity of the biuret method is around 1-10 g protein/l.



Complex of the protein and copper in alkaline environment

## Sources

### Related articles

- Total protein
- Plasma proteins

### References

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